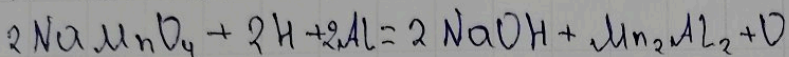
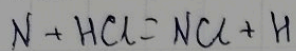
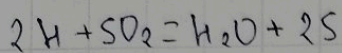
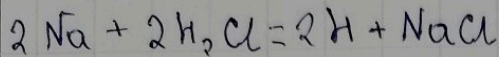


№4



№3

$$3.1) M(\text{C}_8\text{H}_{18}) = M(\text{C}) \cdot 8 + 12 + M(\text{H}) \cdot 18 + 1 = 1042$$

$$\text{C}_8\text{H}_{18} = 2,2 \text{ моль}$$

$$3.2) \text{C}_6\text{H}_6 = M(\text{C}) \cdot 6 + 12 + M(\text{H}) \cdot 6 + 1 = 802$$

$$\text{C}_6\text{H}_6 = 1,2 \text{ моль}$$

№1

1.1) формула:  $\text{MA}_2$  калий йодиді:  $\text{KI}$

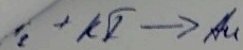
алтын нөубар:  $\text{AuH}$ ; молекуласық формула:  $\text{AuH} + \text{KI} = \text{AuK} + \text{HI}$

1.2) 120 г су; 1, кристалда гидрат

№2

1-жаз:  $\text{C}_2\text{H}_2$ ; 2-жаз:  $\text{H}$

2)



указан way for "размеры" X 9-11.

$$n(Au) = 197 + = e/mole.$$

t = 25°C (MA<sub>2</sub>)

MA<sub>2</sub> = 3 H<sub>2</sub>O

n = 0,7 mole

V = ?

t = 25°C

n<sub>2</sub> MA<sub>2</sub>

V = 15,93 ml H<sub>2</sub>O

m = Ca(OH) = 272

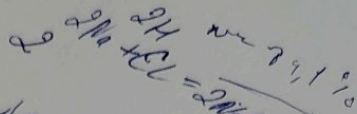
V = (B. 20) = 2,776 n

t = 10°C (MA<sub>2</sub> = 3 H<sub>2</sub>O)

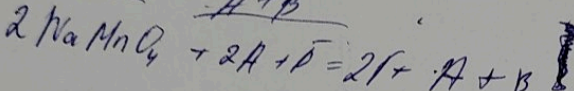
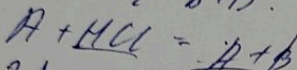
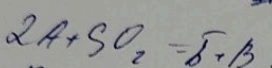
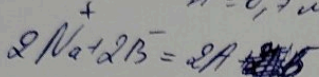
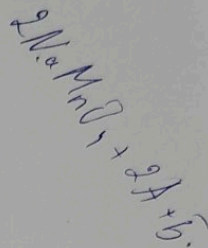
n = 0,7 mole

$$V = \frac{m}{\rho}$$

указан



1/2 CH



CaCl<sub>2</sub>

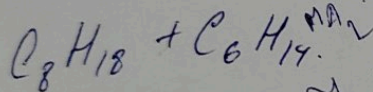
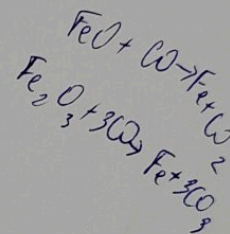
указано



m. eq.

$$V = \frac{m}{\rho} = \frac{27}{1000} = 0,027$$

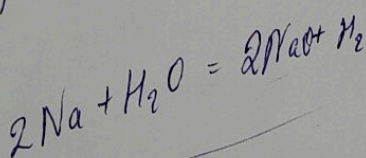
0,027 n



10 H<sub>2</sub>O  
0,515 2/mole

398,95

$$W = \frac{mep \cdot 3}{mep \cdot g} = \frac{8,282}{401}$$



$$m = \rho \cdot V =$$

$$m = \frac{W \cdot m}{100\%} = \frac{21\% \cdot 27}{100\%} = 5,67$$

567

W = ?

$$\frac{5,67}{1000} =$$

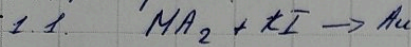
0,00567

0,0206483

2,06483%

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№1 тапсырма



$$M(Au) = 197 + 308,46 = 505,46 \text{ моль}$$

1.2.  $M = 505,45 \text{ г/моль}$

№2 тапсырма

берілгені

$V(H_2O) = 15,93 \text{ мл}$

2.1)

$m(CaOH) = 272$

$V(\text{қанық. р-н}) = 2,576 \text{ л}$   
 $\rho(\text{р-н}) = 1,2 \text{ г/мл}$

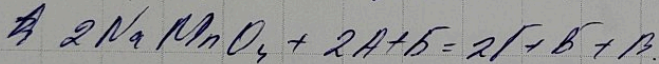
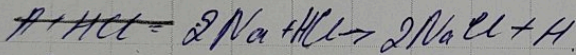
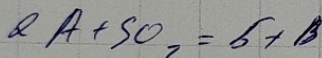
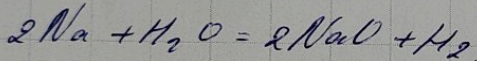
1)  $I$  биіксіз  $CH$

2) сулы сара  
суда р-н  
қолданылған  
архивтегі.

2.2.  $m = \frac{W \cdot m}{100\%} = \frac{21\% \cdot 27}{100\%} = 5,67 \text{ г}$

$V = \frac{m}{\rho} = \frac{5,67 \text{ г}}{1000 \text{ г/л}} = 0,00567 \text{ л}$

№3 тапсырма



№3 тапсырма

3.1.  $W = \frac{m_{ер.3}}{m_{ер.1}} \cdot 100\% = \frac{8,28}{401} \cdot 100\% = 2,06483\%$

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1-мансырма

Берілгені:

Менші

d)

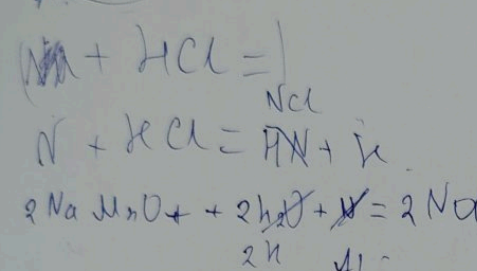
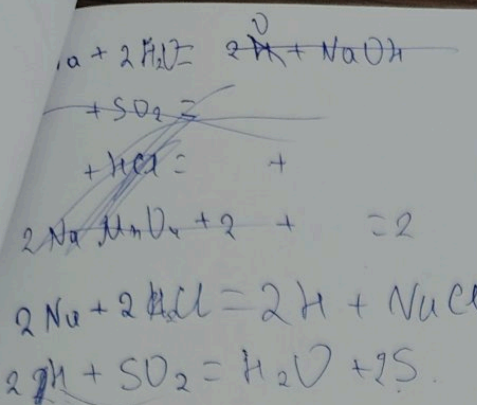
М метал зосы  
лытауы, ұлы  
М метал пидрот  
кауды ерину  
арыбалы МА,  
мүзілеуі.

2-мансырма

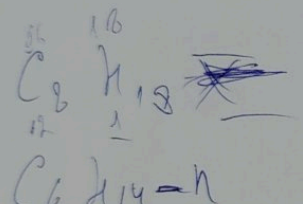
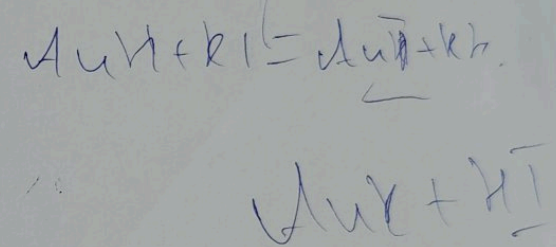
1) Көмір қышқыл газы, көр

3-мансырма

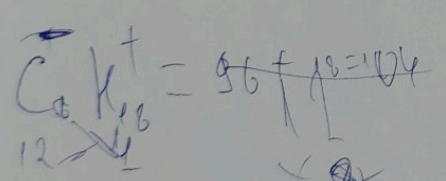
4-мансырма



k.i  
 $\text{MA}_2 + \text{KI} \rightarrow$   
~~+ KI~~  
 KI - хамийногцги.  
 NH<sub>3</sub> - амтагчтай уяан.  
 мөргөлдөөн эрүүлжээ.

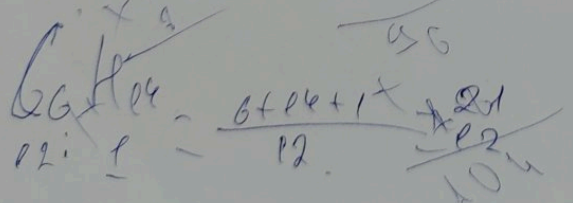


$M(\text{C}_8\text{H}_{18}) = 8 \cdot 12 + 18 \cdot 1 = 96 + 18 = 114$   
 $M(\text{C}_6\text{H}_{12}) = 6 \cdot 12 + 12 \cdot 1 = 72 + 12 = 84$



$\frac{96 + 18}{12} = \frac{9 + 18}{12}$

$\frac{18}{3}$   
 $\frac{26}{24} \cdot 12$   
 $\frac{2}{2} \cdot 22$



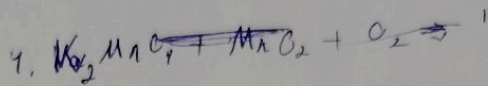
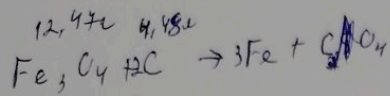
$\frac{72}{6} = 12$

$\frac{21}{22} \cdot 12$

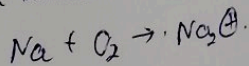
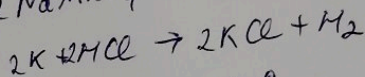
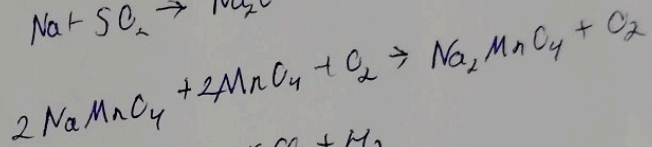
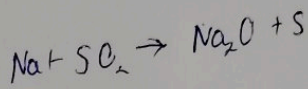
$e_3 O_4 + C$      $MgCl$      $Na$   
                    $BaOH$

~~$MA_2 + AB \rightarrow$~~

X-9-12.



$C_8 H_{18} + C_6 H_{14}$



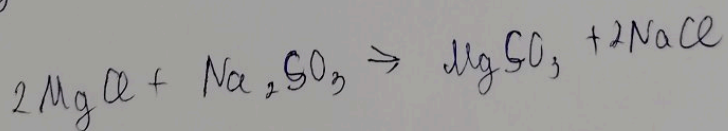
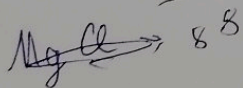
24

16

4

64

24



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2 мап.

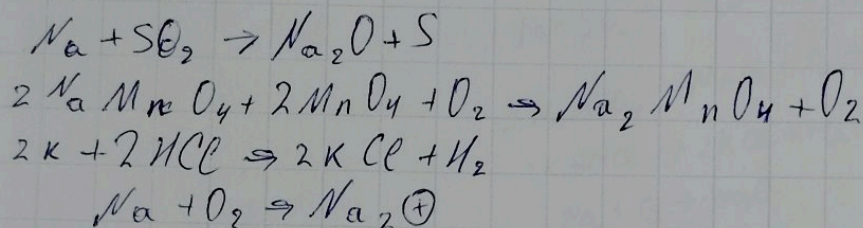
$$15 - 93 \quad 2,576 \quad 0,21 \cdot 27 = 5,67$$

$$x \rightarrow 22,4 \rightarrow 139 \quad 2.21.13.$$

1 мап.

углекислотой, сурьистой роз.

4.1-мап.



3.1 мап.

$$C_8 H_{18} + C_6 = C_{14} H_{31} + H_2 \uparrow$$

$$M_n = 12 \cdot 14 + 31 = 199$$

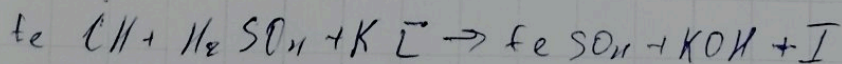
$$W = \frac{8 \cdot 28}{199} = 0,04$$

3.2 мап.

$$M_n = 12 \cdot 10 + 20 = 140$$

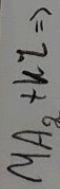
$$W = 140 \cdot 8,410 = 1,1774 \quad C_{10}H_{20} \quad (24,40) \text{ г.мо.}$$

1.1 мап.

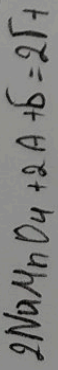
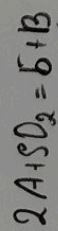
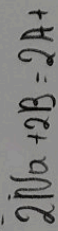


1.2 мап.

X-9-08



масын жазып берүүгө  
у!



31 м

М

15,93 2,576

X 2,4 -> 139

0,21 · 27 = 5,67

0,2 · 1,13

X 31 м

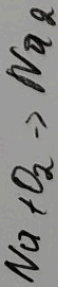
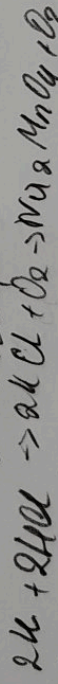
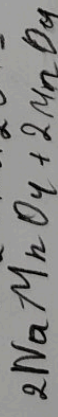
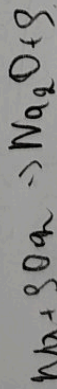
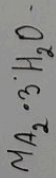
X-узгооруу  $C_2H_{18}$

n-реңеңуу  $C_6H_{14}$

m/бөңүгү - 3,282 - 4014D<sub>2</sub>

реңуу 5500 4D<sub>2</sub>m/м<sub>2</sub>м<sub>2</sub>

+B. H-реңеңуу 42004D<sub>2</sub>m/м<sub>2</sub>м<sub>2</sub>



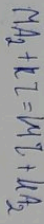
$n = k_1 - k_2$

$m = n \cdot M$

$w = \frac{m}{M}$

$n = \frac{m}{M}$

$kL =$



$MA_2$  эсе

$H_2O \cdot 15,93$

$KOH \cdot 272$

м/бөңүгү 2,576

$C_2H_2$

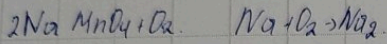
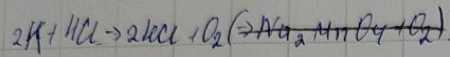
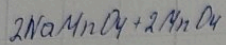
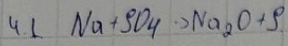
$p = k_1/k_2$

$MA_2 \cdot 3H_2O$



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№4.

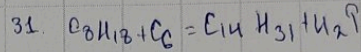


№3.

32.

$$M(n) = 12 \cdot 10 + 20 = 140$$

$$w = 140 \cdot 8 \cdot 4 = 1,1774 \quad \text{C}_{20}\text{H}_{20} \quad (2,4 \cdot 102(100))$$

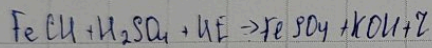


$$m(n) = 12 \cdot 14 + 31 = 199$$

$$v = \frac{8 \cdot 28}{199} = 0,04$$

№1

1.1.



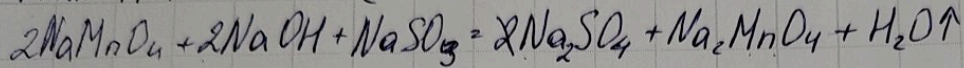
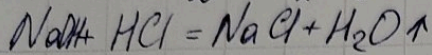
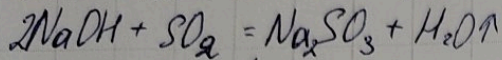
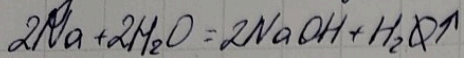
$$1.2. \text{M}_2 \cdot 10\text{H}_2\text{O} \quad \frac{505,45}{10(1 \cdot 2 \cdot 16)} = \frac{505,42}{320} = 1,5795$$

№2.

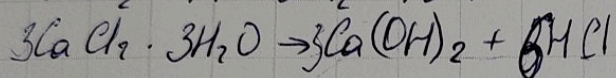
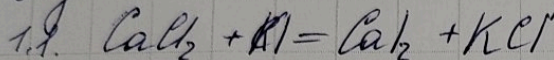
$$15,93 \quad 2,576 \quad 0,2 \cdot 27 = 5,67$$

$$x \quad 2,4 \cdot 5,139 \quad 0,2 \cdot 1,13$$

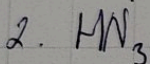
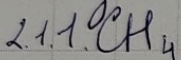
Задача №4. Реакции



Задача №1. Ковиний сахар



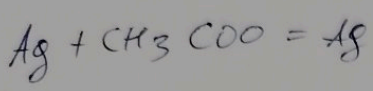
Задача №2



Задача №3

X-9-05.

Уксусная кислота  $CH_3COO$

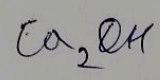


Дано:

$$V(H_2O) 15,93 \text{ мл}$$

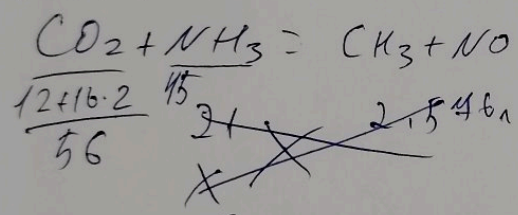
$$m = 242$$

$$V = 2,546 \text{ л}$$



$$P = 1 \text{ атм}$$

$V(O_2)$  ?



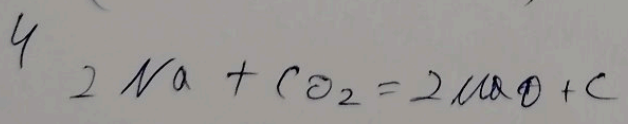
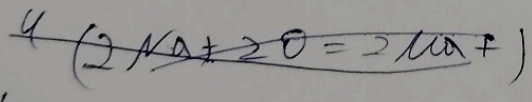
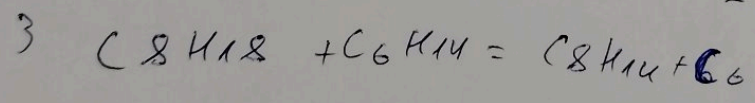
$$\frac{H_2O}{1 \cdot 2 + 16} = 18$$

$$V = ?$$

$$n = \frac{24}{18} = 1,5$$

$$V = 1,5 \cdot 22,4 = 33,6$$

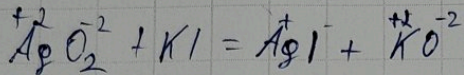
21%



$$n = \frac{m}{M} \quad \rho = \frac{m}{V_m} \quad V = n \cdot V_m$$

1.1

Реакция зәлелешө газға



молекулярная формула:  $Ag_2O_2$

1.2 Дано

$$t = 10^\circ C \rightarrow 25^\circ C$$

газға → 0.7 моль

$$m = 55.22$$

Решение

$$\frac{55.2}{55.2} \cdot 55.2 \cdot X \cdot 25^\circ C = \frac{100 \cdot 25}{55.2} = 45.282$$

$$\frac{29.3}{100} \cdot 100 \cdot X \cdot 10^\circ C = \frac{100 \cdot 10}{29.3} = 34.122$$

n = ?

$$\text{Әуеленше } 0.08958352$$

$$0.06750422$$

$$n = \frac{45.28}{505.45} = 0.08958352$$

$$n = \frac{34.12}{505.45} = 0.06750422$$

2

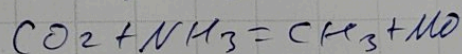
Дано:

$$V (H_2O) = 15.93 \text{ м}$$

$$\text{газға (m)} = 242$$

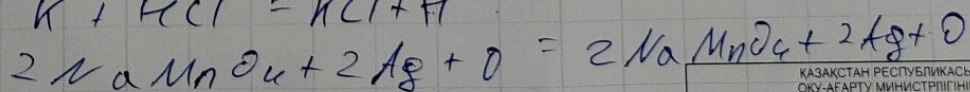
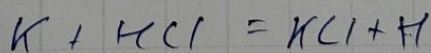
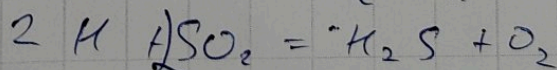
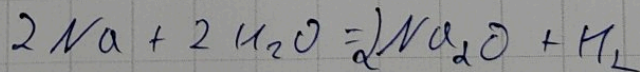
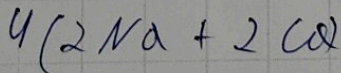
$$V (\text{сеш. газ}) = 2.546 \text{ л}$$

2)  $CO_2, NH_3$



$$\frac{21}{X} \cdot \frac{56}{45} \cdot X = 16.8471$$

3



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2) Задача

0,27 - 27 - 5 - 67

15 - 93: 2,576

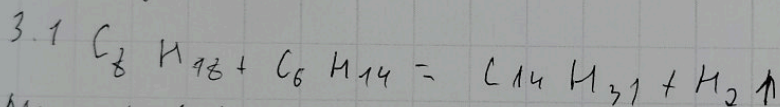
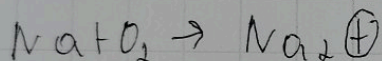
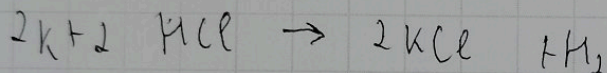
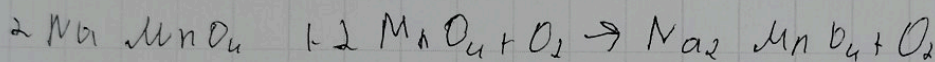
x → 22,4 → 139

2: 21,13

1) Задача

уникальді серпінді газ.

4.1



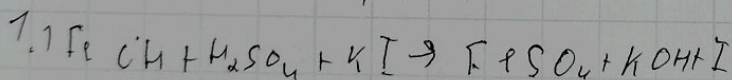
$$M_r = 12 \cdot 14 + 31 = 199$$

$$w = \frac{6 \cdot 28}{199} = 0,04$$

3.2

$$M_r = 12 \cdot 16 + 20 = 140$$

$$w = 140 \cdot 0,410 = 1,1774$$



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Задача №2

Дано:

$$V_1 = 15,93 \text{ м/с}$$

$\rho =$

$$m = 272$$

$$V_2 = 2,576 \text{ м}$$

Задача №1

$$t = 10^\circ\text{C}$$

$$t' = 25^\circ\text{C}$$

$$C_{\text{МА}_2} = 55,2 \text{ ккал} = 0,7 \text{ моль}$$

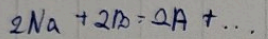
$$\rho_1 = 55,2 \text{ ккал} \cdot 100 \text{ г}$$

$$\rho_2 = 29,3 \text{ ккал} \cdot 100 \text{ г}$$

$$t = 25^\circ\text{C}$$

$$t' = 10^\circ\text{C}$$

Задача №4



$$2 \cdot 23 + 2 = 48 = T_i$$

Задача №3

$$T = 40^\circ\text{C}$$

$$T_{\text{ср}} = 2,25 \cdot 5500 \text{ кДж/моль}$$

$$T = 4200 \text{ кДж/моль}$$

Решение:

$$1) 500 \cdot 5,45 - 0,7 = 504,75$$

$$0,7 \cdot 170 = 119$$

$$505,45 - 119 = 386,45$$

Решение:

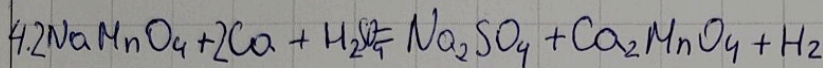
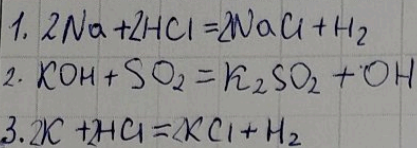
$$1) 12 \cdot 8 + 1 \cdot 18 = 114$$

$$2) 5500 \text{ кДж/моль} - 4200 \text{ кДж/моль} = 1300 \text{ кДж/моль}$$

моль.

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N 4.1



N 3.1

Дано  
 $m = 8,282$   
 $a = 401,2 \text{ Dn}$   
 $Q(\text{C}_8\text{H}_{18}) = 5500$   
 $Q(\text{C}_6\text{H}_6) = 4200$

Решение

$$\text{C}_8\text{H}_{18} + \frac{\text{C}_6\text{H}_6}{88} = \frac{\text{C}_6\text{H}_6}{78}$$

$$M(\text{C}_6\text{H}_6) = 72 + 18 = 114 \text{ моль}$$

$$M(\text{C}_8\text{H}_8) = 72 + 6 = 78 \text{ моль}$$

$$\omega = \frac{4200 \cdot 78}{88} = 3276\%$$

$\omega(x) = ?$   
 $M = ?$

N 3.2

Дано  
 $\text{C}_8\text{H}_{18}$   
 $\text{C}_6\text{H}_{14}$   
 $\omega(\text{C}) = 84,1\%$

Решение

$$84,1 - 100\% = 15,9\% - \gamma$$

$$\frac{\text{C}_8\text{H}_{18}}{114} + \frac{\text{C}_6\text{H}_6}{78} =$$

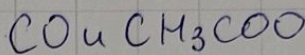
$M(x) = ?$   
 $M(y) = ?$

$$M(\text{C}_8\text{H}_{18}) = 96 + 18 = 114 \text{ моль}$$

$$M(\text{C}_6\text{H}_6) = 72 + 6 = 78 \text{ моль}$$

$$\omega = \frac{5500 \cdot 78}{114} = 4250\%$$

N 2. 1.



N 2. 2. 1

Дано  
 $0 - 21\%$   
 $V(\text{H}_2\text{O}) = 15,93\%$   
 $m(\text{Ca}(\text{OH})_2) = 27$   
 $V = 2,576\%$

Решение

$$\text{H}_2\text{O} + \frac{\text{Ca}(\text{OH})_2}{74} = \frac{\text{CO}}{28} + \text{CH}_3\text{COO}$$

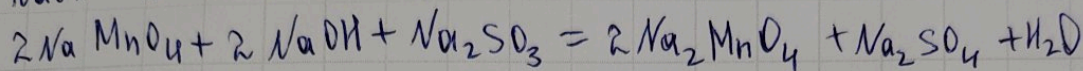
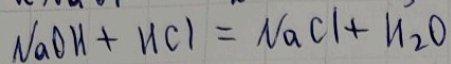
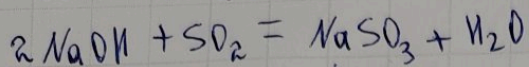
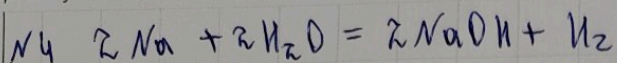
$$M(\text{CO}) = 12 + 16 = 28 \text{ моль}$$

$$M(\text{Ca}(\text{OH})_2) = 40 + 17 \cdot 2 = 74 \text{ моль}$$

$$V = \frac{m}{M_n} = \frac{27}{74} = 0,481$$

$$\omega = \frac{27 \cdot 28}{74} = 10,2\%$$

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N3 3.1. Массы ұзектаны ( $C_8H_{18}$ )  $m_1 = 5,70г$

$w_1 = 68,8\%$

Массы n-гексаны ( $C_6H_{14}$ )  $m_2 = 2,58г$

$w_2 = 31,2\%$

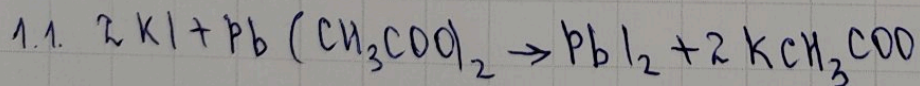
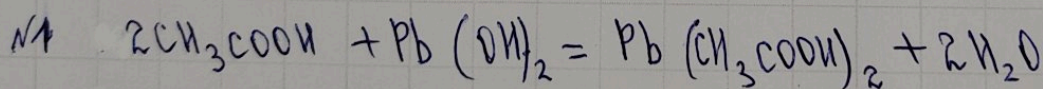
Маллыя дала ұзектаны  $62,5\%$

Маллыя дала n-гексаны  $37,5\%$

3.2.

N2 1) два газа:  $H_2$  и  $CO_2$

2)  $V(\text{воздуха}) = 23,6л$



1.2.